Electronegativity Scales

The Pauling scale is based on the difference between the bond energy for a molecule AB and the average bond energy for the molecules AA and BB; that is, it measures the extent to which the bond AB is stronger than one might expect given the average strength of an AA bond and a BB bond. Hydrogen is assigned a Pauling electronegativity of 2.0 and all other values are referenced to this.

In the Allen scale, an element’s electronegativity is equal to its average valence electron energy.