Electronegativity using the Allen scale

Group →	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
↓ Period																		
1	H 2.300																	He 4.160
2	Li 0.912	Be 1.576											B 2.051	C 2.544	N 3.066	O 3.610	F 4.193	Ne 4.789
3	Na 0.869	Mg 1.293											Al 1.613	Si 1.916	P 2.253	S 2.589	Cl 2.869	Ar 3.242
4	K 0.734	Ca 1.034	Sc 1.19	Ti 1.38	V 1.53	Cr 1.65	Mn 1.75	Fe 1.80	Co 1.84	Ni 1.88	Cu 1.85	Zn 1.59	Ga 1.756	Ge 1.994	As 2.211	Se 2.434	Br 2.685	Kr 2.966
5	Rb 0.706	Sr 0.963	Y 1.12	Zr 1.32	Nb 1.41	Mo 1.47	Tc 1.51	Ru 1.54	Rh 1.56	Pd 1.59	Ag 1.87	Cd 1.52	In 1.656	Sn 1.824	Sb 1.984	Te 2.158	I 2.359	Xe 2.582
6	Cs 0.659	Ba 0.881	Lu 1.09	Hf 1.16	Ta 1.34	W 1.47	Re 1.60	Os 1.65	Ir 1.68	Pt 1.72	Au 1.92	Hg 1.76	Tl 1.789	Pb 1.854	Bi 2.01	Po 2.19	At 2.39	Rn 2.60
7	Fr 0.67	Ra 0.89																

See also: Electronegativities of the elements (data page)

Packing Possibilities Assuming Anion is Larger Than Cation[a]

anion's lattice	cation's hole	Coordination Number (max) cation	Coordination Number (max) anion	r_{cation}/r_{anion}	base stoichiometry[b]	other stoichiometries[c]
simple cubic	cubic	8	8	0.732 – 0.999	1:1	1:2, 1:4
face-centered	octahedral	6	6	0.414 – 0.732	1:1	1:2, 2:3, 1:3
face-centered	tetrahedral	4	8	0.225 – 0.414	2:1	1:1, 1:2, 3:2

[a] when cations are larger than anions, simply reverse their roles in the table

[b] ratio is cation:anion assuming all holes are filled

[c] assuming that some fraction of the holes remain unfilled