

## Naming Simple Compounds

### Set 1

N: atomic nitrogen

N<sub>2</sub>: molecular nitrogen

NO: nitrogen monoxide

NO<sub>2</sub>: nitrogen dioxide

N<sub>2</sub>O: dinitrogen monoxide

N<sub>2</sub>O<sub>5</sub>: dinitrogen pentoxide

S<sub>4</sub>N<sub>3</sub>: tetrasulfur trinitride

Cl<sub>2</sub>O: dichlorine monoxide

PO<sub>3</sub>: phosphorous trioxide

Cl<sub>2</sub>O<sub>7</sub>: dichlorine heptoxide

H<sub>2</sub>S: dihydrogen sulfide

### Set 2

FeCl<sub>2</sub>: iron(II) chloride

FeCl<sub>3</sub>: iron(III) chloride

ZnCl<sub>2</sub>: zinc chloride

Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>: iron(III) sulfate

FeO: iron(II) oxide

Fe<sub>2</sub>O<sub>3</sub>: iron(III) oxide

Fe<sub>3</sub>O<sub>4</sub>: iron (II, III) oxide

KClO<sub>4</sub>: potassium perchlorate

NaClO<sub>3</sub>: sodium chlorate

Fe(ClO<sub>2</sub>)<sub>2</sub>: iron (II) chlorite

Fe(ClO)<sub>3</sub>: iron (III) hypochlorite