

Name the following...

CBr_4 : *carbon tetrabromide* (covalent compound)

SrCl_2 : *strontium chloride* (ionic compound; strontium has just a single common cation with a charge of +2, so there is no need to indicate the charge)

$\text{Co}(\text{NO}_3)_2$: *cobalt(II) nitrate* (ionic compound; cobalt has common cations with charges of +2 and of +3, so we must include the cation's charge in the name to avoid uncertainty; NO_3^- is nitrate, one of the common polyatomic anions)

KCN : *potassium cyanide* (ionic compound; CN^- is cyanide, one of the common polyatomic anions)

Cl_2O : *dichlorine monoxide* (covalent compound)

Write formulas for the following...

magnesium fluoride: *MgF_2* (ionic compound; magnesium has a single common cation with a charge of +2 and fluorine has a single common anion with a charge of -1)

copper(I) chloride: *CuCl* (ionic compound; copper has two common cations with charges of +1 and of +2, but here the inclusion of the Roman numeral I tell us that it is +1; chlorine has a single common anion with a charge of -1)

potassium carbonate: *K_2CO_3* (ionic compound; potassium has a single common cation with a charge of +1 and carbonate is the polyatomic anion CO_3^{2-})

tetranitrogen oxide: *N_4O* (covalent compound)

sulfur trioxide: *SO_3* (covalent compound)